

Covering Chap. 12 & 13 in this order (very approximate):

- 12.1 Electronic config of monatomic ions
 - 12.2 Lewis struc of monatomic ions and ionic compound
 - 12.3 Covalent bonds of simple molecules
 - 12.4 Electronegativity trend, Polarity of bonds
 - 13.1 Drawing Lewis structures of NH_3 , CCl_4 , H_2O , ClO^- , ammonium ion, sodium hydroxide, magnesium hydroxide, magnesium chloride, oxygen gas, nitrogen gas.
-

- 13.2 Electron-Pair Repulsion of water, ammonia, carbon tetrachloride
- 13.3 Molecular Geometry of water, ammonia, carbon tetrachloride

Repeat 13.1-13.3 (incorporating 12.5 & 12.6) and add in 13.4 & 13.5.

Drawing Lewis structures of more complex molecules and polyatomic ions & predict molecular geometry & molecular polarity (O_2 , N_2 , CO_2 , C_2H_4 , C_2H_2 , HCN , formaldehyde H_2CO , hydroxide, ammonium ion, nitrite, carbonate, nitrate, acetate)

- 12.7 Exceptions to the Octet Rule:
 - Odd # of electron (NO , NO_2)
 - More than Octet (PF_5 , SF_6)
 - Less than Octet (remember Be and B: BeF_2 and BF_3)